

# Classifying Matter

## KEY IDEAS

As you read this section, keep these questions in mind:

- How can matter be classified?
- Why are elements and compounds considered pure substances?
- How are elements and compounds different from mixtures?

## What Is Matter?

This book, a pencil, a piece of paper, and even you are made of matter. In fact, anything that you can hold or touch is made of matter. **Matter** is anything that has both mass and volume. *Mass* consists of all the particles that make up an object. *Volume* is how much space the object occupies. ✓

Some things you cannot see are matter. For example, you cannot see air, but because it has both mass and volume, it is matter. Some things you can see, however, are not matter. For example, light is not matter even though you can see it. Why isn't light matter? It does not have mass or volume.

## MATTER AND CHEMISTRY

*Chemistry* is the branch of science that studies matter and its changes. Chemists study the particles that make up matter. They also study what happens to these particles as they interact with one another. Chemists often use the knowledge they gain to develop new products for consumers such as you. Almost everything that you use in daily life involves chemistry. This includes foods, soaps, clothing, and compact discs (CDs).



## READING TOOLBOX

**Compare** After you read this section, make a Venn diagram to show the similarities and differences between elements, compounds, and mixtures.



## READING CHECK

**1. Identify** What two properties does all matter have?

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## LOOKING CLOSER

**2. Explain** What can you see in this picture that is not matter? Explain your answer.

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**SECTION 1** Classifying Matter *continued*

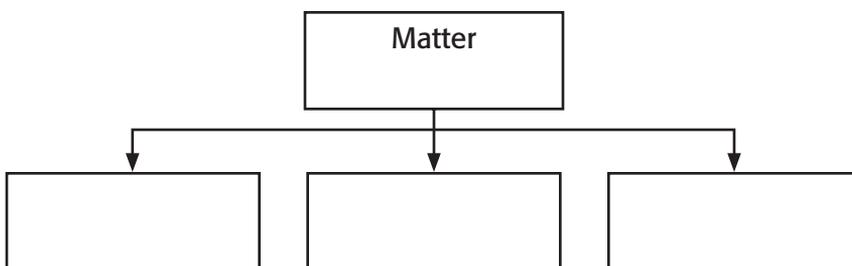
**CLASSIFYING MATTER**

To *classify* something is to place it into a specific group. All members in the group share certain features. For example, the CDs in a music store are classified into groups such as rock, classical, and jazz. Like CDs, matter is classified into different groups.

Matter can be classified in several ways. One of the most common ways scientists group matter is based on what makes it up. Every sample of matter is either an element, a compound, or a mixture. Gold and oxygen are examples of elements. Water and sugar are examples of compounds. Vinegar and brass are examples of mixtures.

**LOOKING CLOSER**

**3. Complete** Fill in the graphic organizer to show the three main groups of matter.

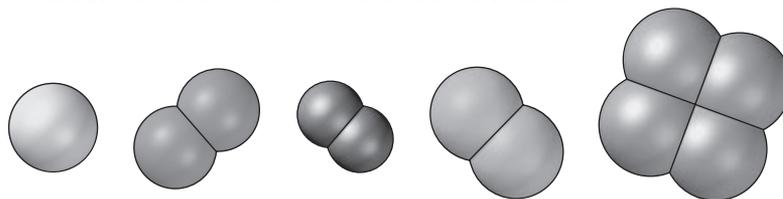


**What Is an Element?**

Gold and oxygen are classified as elements. An **element** is a substance that cannot be separated or broken down into simpler substances. An element is made of only one kind of atom. An **atom** is the smallest unit of an element that has the chemical properties of that element.

Some elements exist as single atoms. Helium, an element often used to fill balloons, exists as a single atom. Other elements exist as combinations of more than one atom known as molecules. A **molecule** is a combination of two or more atoms that are combined in a definite ratio. For example, a molecule of water has two hydrogen atoms and one oxygen atom. The ratio of hydrogen atoms to oxygen atoms in water is always 2:1.

Each of these molecules is made up of atoms of the same element. Many molecules are made of two or more different elements.



Neon      Oxygen      Hydrogen      Chlorine      Phosphorus

**Critical Thinking**

**4. Identify Relationships**

How are atoms and molecules related?

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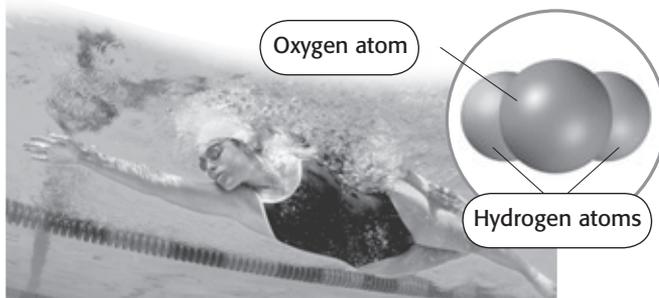
**SECTION 1** Classifying Matter *continued***ELEMENT SYMBOLS**

Each element is represented by a one- or two-letter symbol. The symbol is either a single capital letter or a capital letter and a lowercase letter. For example, the symbol for carbon is C and the symbol for copper is Cu. All the elements and their symbols are listed in the periodic table.

**What Is a Compound?**

Water and sugar are examples of compounds. A **compound** is a substance made up of atoms of different elements that are chemically combined. When elements combine to form a compound, they always combine in the same ratio. For example, a water molecule is always made of two hydrogen atoms and one oxygen atom.

Water is a compound. Each water molecule is made up of two hydrogen atoms and one oxygen atom.

**UNIQUE PROPERTIES**

The properties of a compound are different from the properties of the elements that make up the compound. For example, both hydrogen and oxygen are gases at room temperature. However, hydrogen and oxygen combine chemically to form the compound water, which is liquid at room temperature.

**CHEMICAL FORMULAS**

A *chemical formula* shows how many atoms of each element are in a unit of a substance. The number of atoms of each element appears as a *subscript* after the symbol. For example, a molecule of  $O_2$  has two oxygen atoms.

Numbers placed in front of a chemical formula show how many molecules of a substance that you have. One molecule of table sugar is written as  $C_{12}H_{22}O_{11}$ . Therefore, three molecules of sugar are written as  $3C_{12}H_{22}O_{11}$ .

**Critical Thinking**

**5. Explain** A particular molecule has two hydrogen atoms and two oxygen atoms. How do you know that this is not a molecule of water?

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**Math Skills**

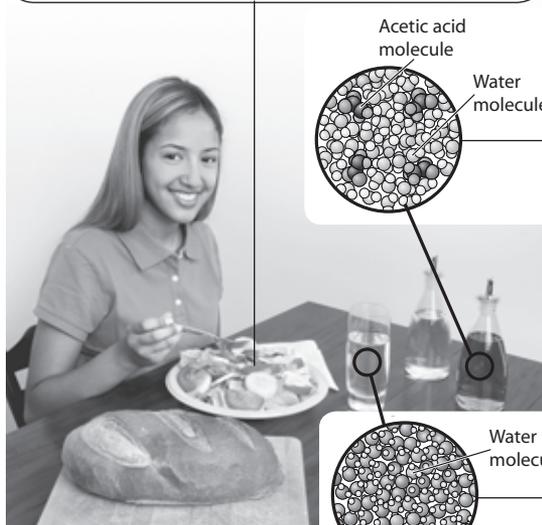
**6. Calculate** How many total atoms are in three molecules of table sugar?

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**SECTION 1** Classifying Matter *continued*

A tossed salad is an example of a *heterogeneous mixture*. The substances in a heterogeneous mixture are not distributed, or spread, evenly throughout the mixture. Different parts of the mixture may not be the same. Some parts of the salad, for example, may have more tomatoes or carrots than other parts of the salad.



The vinegar used on a salad is a *homogeneous mixture*. The components of a homogeneous mixture are distributed evenly. This means that every part of the mixture is the same. In vinegar, molecules of acetic acid are spread evenly among the molecules of water

Water is a *pure substance*. It has a fixed composition and definite properties.

**LOOKING CLOSER**

**10. Compare** How do homogeneous and heterogeneous mixtures differ?

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**MISCIBLE AND IMMISCIBLE**

Gasoline is a homogeneous mixture made up of many different liquids. The liquids in gasoline are *miscible*, which means that they can be mixed and will stay mixed.

If you shake up a mixture of oil and water, they will not mix well. The oil will rise to the top of the container and the water will stay at the bottom. Oil and water are *immiscible*, which means they do not mix or will not stay mixed.

**Critical Thinking**

**11. Apply Concepts** Generally, you must stir a can of paint before you use it. Is paint miscible or immiscible? Explain your answer.

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**MAKING MIXTURES**

A mixture can be made up of solids, liquids, or gases. It can even be made up of all three. For example, a carbonated drink is a homogeneous mixture made up of water, sugar, flavorings, and carbon dioxide gas. Some other examples of mixtures are shown below.

Example	A mixture of...
Air	gases
Sand	solids
Gasoline	liquids
Fog	gases and liquid
Seawater	solids, liquids, and gases

# Section 1 Review

## SECTION VOCABULARY

**atom** the smallest unit of an element that maintains the chemical properties of that element

**compound** a substance made up of atoms of two or more different elements joined by chemical bonds

**element** a substance that cannot be separated or broken down into simpler substances by chemical means; all atoms of an element have the same atomic number

**matter** anything that has mass and takes up space

**mixture** a combination of two or more substances that are not chemically combined

**molecule** a group of atoms that are held together by chemical forces; a molecule is the smallest unit of matter that can exist by itself and retain all of a substance's chemical properties

**pure substance** a sample of matter, either a single element or a single compound, that has definite chemical and physical properties

**1. Explain** Why is Co an element, but CO a compound? (Hint: Use the periodic table in the back of the book to help you.)

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**2. Compare** How are compounds and mixtures alike? How are they different?

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**3. Complete** Fill in the concept map below to show the relationship between the following: compound, element, matter, mixture, pure substance.

